

Cite this: *Dalton Trans.*, 2022, **51**, 5082

Synthesis of substituted (N,C) and (N,C,C) Au(III) complexes: the influence of sterics and electronics on cyclometalation reactions†

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Cyclometalated Au(III) complexes are of interest due to their catalytic, medicinal, and photophysical properties. Herein, we describe the synthesis of derivatives of the type (N,C)Au(OAc^F)₂ (OAc^F = trifluoroacetate) and (N,C,C)AuOAc^F by a cyclometalation route, where (N,C) and (N,C,C) are chelating 2-arylpyridine ligands. The scope of the synthesis is explored by substituting the 2-arylpyridine core with electron donor or acceptor substituents at one or both rings. Notably, a variety of functionalized Au(III) complexes can be obtained in one step from the corresponding ligand and Au(OAc)₃, eliminating the need for organomercury intermediates, which is commonly reported for similar syntheses. The influence of substituents in the ligand backbone on the resulting complexes was assessed using DFT calculations, ¹⁵N NMR spectroscopy and single-crystal X-ray diffraction analysis. A correlation between the electronic properties of the (N,C) ligands and their ability to undergo cyclometalation was found from experimental studies combined with natural charge analysis, suggesting the cyclometalation at Au(III) to take place *via* an electrophilic aromatic substitution-type mechanism. The formation of Au(III) pincer complexes from tridentate (N,C,C) ligands was investigated by synthesis and DFT calculations, in order to assess the feasibility of C(sp³)-H bond activation as a synthetic pathway to (N,C,C) cyclometalated Au(III) complexes. It was found that C(sp³)-H bond activation is feasible for ligands containing different alkyl groups (isopropyl and ethyl), although the C-H activation is less energetically favored compared to a ligand containing *tert*-butyl groups.

Received 6th February 2022

Accepted 4th March 2022

DOI: 10.1039/d2dt00371f

rsc.li/dalton

Introduction

(N,C)-Cyclometalated Au(III) complexes^{1–4} have found application within catalysis^{5–9} and medicine^{10–15} and have also gained attention for their photophysical properties.^{3,4,16–18} While cyclometalated Au(III) dichloro complexes have been known since the 1980s,¹ corresponding acetate and trifluoroacetate complexes are less studied. Due to the labile nature of carboxylate ligands, the reactivity of such complexes in *e.g.* ligand

exchange reactions is higher than for their chloride analogues.² Trifluoroacetate ligands are especially labile, making Au(III) trifluoroacetate complexes attractive catalyst candidates,^{5,7} as well as useful intermediates for the synthesis of other Au(III) complexes.^{19–22} In 2012 the first di(trifluoroacetate) complex of Au(III) with a chelating 2-arylpyridine ligand was reported by our group (**2a-Au(OAc^F)₂**, Fig. 1).¹⁹ This complex was conveniently synthesized by reacting tpy (2-(*p*-tolyl)-pyridine, **1a**) with Au(OAc)₃ in a 1 : 1 mixture of HOAc^F (trifluoroacetic acid) and water, using microwave-heating. This was a further development of a protocol for the synthesis of (N,C) Au(III) dichloro complexes, such as **2a-AuCl₂**²³ (Fig. 1). The reactivity of **2a-Au(OAc^F)₂** towards organolithium and Grignard reagents,^{19,20} ethylene,^{24,25} higher alkenes²⁶ and acetylene⁷ has since then been explored. Subsequent to the initial findings, only few examples of trifluoroacetate complexes of Au(III) with 2-arylpyridine ligands have been reported.^{27–29} The scope of the microwave-assisted cyclometalation protocol has not been systematically investigated, although microwave-mediated synthesis of Au(III) complexes has had an increase in popularity in the last decade.^{3,30–36}

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† Electronic supplementary information (ESI) available. CCDC 2115512, 2122527, 2085151, 2086346, 2126159, 2126283, 2122284, 2111917, 2105655, 2086931, 2126114, 2130186, 2126097 and 2114274. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d2dt00371f

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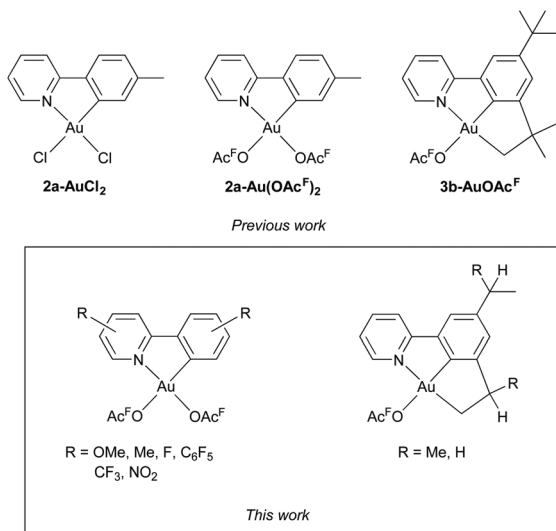


Fig. 1 Top: previously synthesized Au(III) complexes by our group: the (N,C) complexes **2a-AuCl₂**²³ and **2a-Au(OAc^F)₂**¹⁹ and the (N,C,C) complex **3b-AuOAc^F**.⁸ Bottom: (N,C) and (N,C,C) Au(III) trifluoroacetate complexes studied herein.

(N,C,C) Au(III) pincer complexes have been reported to be more stable than (N,C) systems towards protolytic decomposition,⁸ and also possess interesting luminescence properties compared to (C,N,C) and (N,C) Au(III) complexes.¹⁶ For Au(III), the majority of reported (N,C,C) pincer ligands are derived from 2-([1,1'-biphenyl]-3-yl)pyridine (Fig. 2).^{2,16,18} While cyclometalation through C(sp³)-H bond activation is common for *e.g.* Pd(II),^{37–40} reports on C(sp³)-H bond activation in the synthesis of Au(III) complexes are scarce, and only a few examples have been reported.^{41–43} In 2018, we reported a rare example of C(sp³)-H bond activation for Au(III) in the synthesis of an (N,C,C)-cyclometalated Au(III) trifluoroacetate complex **3b-AuOAc^F** from 2-(3,5-di-*tert*-butylphenyl)pyridine (**1b**) (Fig. 1).⁸ The formation of complex **3b-AuOAc^F** from **1b** was studied in detail by DFT calculations, showing that the steric bulk of the *tert*-butyl substituent promotes the C(sp³)-H activation by destabilising the intermediate (N,C) complex.

To understand and develop the chemistry of organometallic Au(III) complexes, robust synthesis protocols are of high importance. Varying the functionalization of the ligand backbone in coordination compounds is of interest as this can have a significant impact on their catalytic,^{44–47} photophysical,^{2,16–18,48–55} magnetic,^{56,57} electrochemical^{58–62} and biological^{63–65}

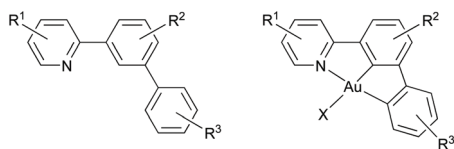


Fig. 2 Ligand for (N,C,C) Au(III) pincer formation derived from 2-([1,1'-biphenyl]-3-yl)pyridine (left) and the corresponding Au(III) complex that possess two Au-C(sp²) bonds (right).

properties. Furthermore, it allows for the evaluation of the robustness of the metalation protocol. Therefore, we wanted to explore the possibility to synthesize functionalized derivatives of **2a-Au(OAc^F)₂** and **3b-AuOAc^F** (see general structures in Fig. 1). Variation of the ancillary 2-arylpyridine ligand is easily implemented, particularly through cross-coupling reactions, and by this a series of new ligands for Au(III) is readily available. In addition, cyclometalation as a strategy for the synthesis of (N,C) and (N,C,C) Au(III) complexes remains a somewhat underdeveloped field. More traditional approaches, such as transmetalation of the corresponding organomercury compounds, are frequently being reported for the synthesis of cyclometalated Au(III) complexes.^{1,2,18,65–68} Although efficient, this method suffers from the toxicity of mercury, creating a need to investigate and further develop alternative synthesis methods. We herein present the synthesis and characterization of a series of (N,C) and (N,C,C) Au(III) complexes. All complexes were conveniently prepared by microwave-heating using Au(OAc)₃ and the corresponding 2-arylpyridine ligand, with electron-donating or -withdrawing substituents at one or both rings. The formation and bonding properties of these complexes were assessed by DFT calculations,¹⁵ NMR spectroscopy and single-crystal X-ray diffraction analysis, in order to address any substituent effects.

Results and discussion

Ligand synthesis

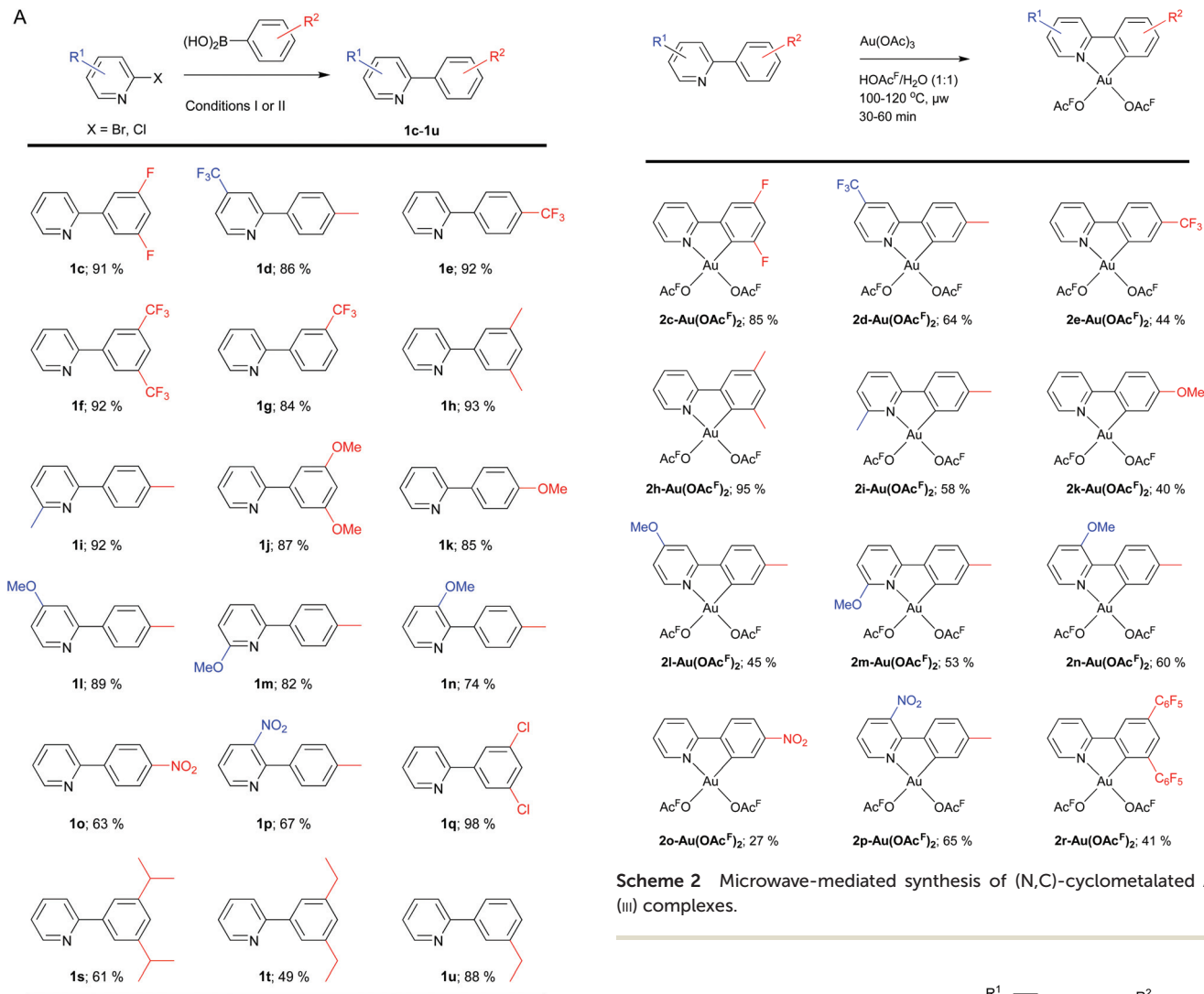
The ligands, substituted 2-arylpyridines (**1c–1u**), were readily available through the Suzuki–Miyaura cross-coupling of suitable 2-halogenated pyridine derivatives and arylboronic acids (Scheme 1A), using reaction conditions reported in the literature.^{8,69,70} Ligand **1r** was prepared from **1q** (Scheme 1B) according to a method developed by Fagnou and co-workers for the installation of pentafluorophenyl groups through Pd-catalysed C–H activation.⁷¹

Synthesis of (N,C)-cyclometalated Au(III) di(trifluoroacetate) complexes

Having in hand a wide variety of potential ligands, we investigated their ability to cyclometalate at Au(III) under the same reaction conditions as utilized for the synthesis of **2a-Au(OAc^F)₂** (Scheme 2).²³

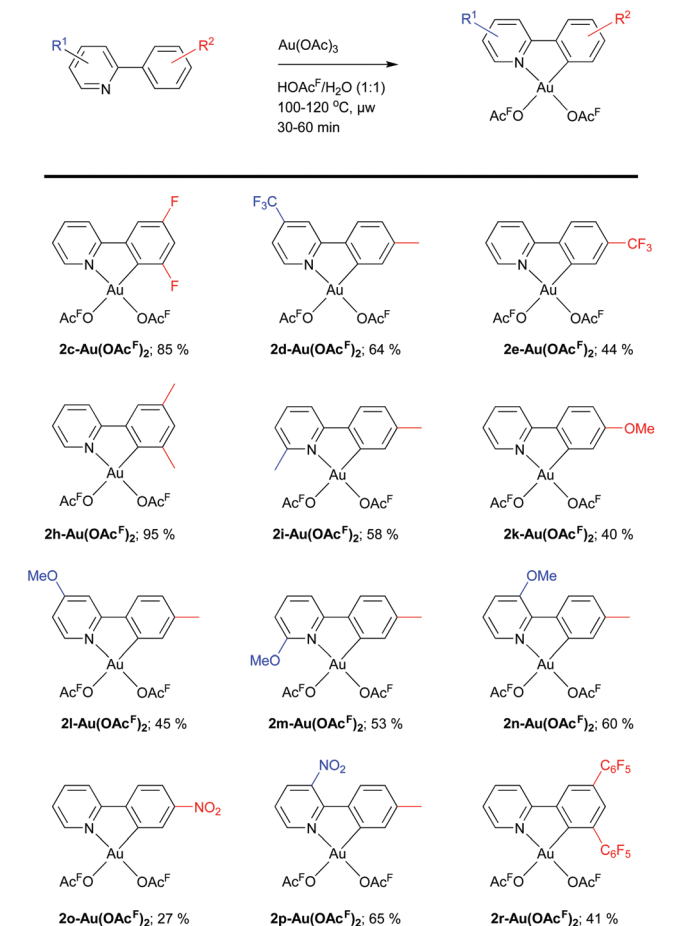
The Au(III) complexes were obtained in yields ranging from 27 to 95%, and both electron-withdrawing (nitro, trifluoromethyl) and electron-donating (methyl, methoxy) substituents were tolerated. They were characterized by multinuclear NMR spectroscopy (¹H, ¹³C, ¹⁹F and ¹⁵N NMR), MS, elemental analysis and single-crystal X-ray diffraction analysis. For Au(III) complexes carrying substituents on the pyridine ring, no clear trends in yields were observed. For complexes with substituents on the phenyl ring, certain trends in yields were found. For evaluation of the experimental observations, the carbon that undergoes cyclometalation in the protonated ligands (C2', see Scheme 3) was investigated by natural charge analysis. In



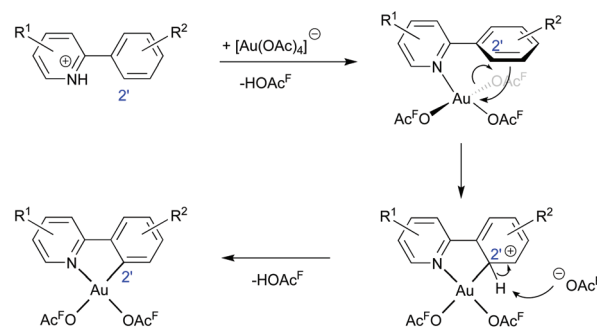


Scheme 1 (A) Synthesis of 2-arylpyridine ligands. Conditions I = Pd(OAc)₂, PPh₃, K₃PO₄, H₂O, *n*-PrOH.⁸ Conditions II = Pd₂dba₃, HBF₄·P(*t*-Bu)₃, KF·2H₂O, THF.^{69,70} (B) Synthesis of ligand **1r** from **1q**.

the following section, the natural charge of C2' in the ligands are discussed relative to the charge of C2' in the 2-phenylpyridinium cation, **ppy-H⁺** ($\Delta C2' = 0$) (for full details, see ESI†). The dimethyl-substituted complex **2h-Au(OAc^F)₂** was obtained in high yield (95%) like complex **2a-Au(OAc^F)₂** (94%).¹⁹ Complexes carrying either significantly electron-withdrawing (**2e-Au(OAc^F)₂**, **2o-Au(OAc^F)₂** and **2r-Au(OAc^F)₂**) or electron-donating (**2k-Au(OAc^F)₂**) groups were obtained in lower yields. The difluoro-substituted complex **2c-Au(OAc^F)₂** was obtained



Scheme 2 Microwave-mediated synthesis of (N,C)-cyclometalated Au(III) complexes.



Scheme 3 Postulated mechanism for the formation of cyclometalated Au(III) complexes. The carbon (C2') that undergoes the C(sp²)-H bond activation and binds to gold in the final product is indicated.

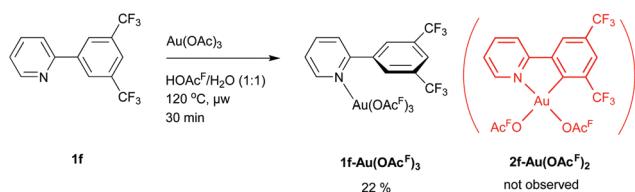
in 85% yield, supporting an electrophilic aromatic substitution-type of cyclometalation mechanism (Scheme 3). This mechanism is generally preferred for the formation of cyclometalated complexes of late transition metals.⁷²⁻⁷⁶ Despite being inductively electron-withdrawing, fluorine groups are *ortho/para*-directing and activating substituents in aromatic electrophilic substitution reactions, causing the C(sp²)-H acti-



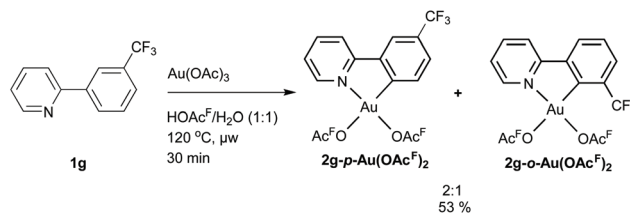
vation to proceed more easily at ligand **1c** compared to *e.g.* **1e**, **1o** and **1r**.

Neither di(trifluoromethyl)- nor dimethoxy-substituted ligands **1f** and **1j** provided the desired cyclometalated products. The reaction of **1j** and Au(OAc)₃ furnished a multitude of species as seen in the ¹H NMR spectrum of the crude product (Fig. S89, ESI[†]), and no single, clean compound could be isolated. It was however possible to obtain crystals suitable for single-crystal XRD, showing an unusual dinuclear M₂L₃-type complex (**2j₃Au₂(OAc^F)₂**), see Fig. S146 and Fig. S147, ESI[†]). Having said this, we do not believe this to be the main reaction product, as no other characterization data can substantiate this. We surmise that the ligand is too electron rich and reactive to yield clean formation of a Au(III) complex under the investigated reaction conditions. This assumption is further supported by the calculated natural charge for C2' in **1j-H⁺** ($\Delta C2' = -0.120$). The natural charge for C2' in **1j-H⁺** is significantly larger than the ones found for ligands **1c-H⁺** ($\Delta C2' = -0.084$) and **1h-H⁺** ($\Delta C2' = -0.008$), which both have activating substituents in the 3'- and 5'-positions. For di(trifluoromethyl)-substituted **1f**, the *N*-coordinated adduct **1f-Au(OAc^F)₃** (which is a likely precursor for the cyclometalation step) was isolated (Scheme 4).

The failure of obtaining cyclometalated **2f-Au(OAc^F)₂** can be related with the poor electrophilicity of **1f** ($\Delta C2' = +0.061$ for **1f-H⁺**). It was previously reported by our group that pincer complex **3b-AuOAc^F** (derived from the sterically encumbered and electron rich ligand **1b**) proceeds *via* the corresponding (N,C)-cyclometalated complex **2b-Au(OAc^F)₂** (see below).⁸ As the trifluoromethyl group is smaller than the *tert*-butyl group,^{77,78} the formation of **2f-Au(OAc^F)₂** should be feasible from a steric point-of-view. If the Au-C bond formation takes place by electrophilic C(sp²)-H bond activation (Scheme 3), (a) strongly electron-withdrawing substituent(s) in the aryl ring of the ligand might impede the reaction. Reaction of the mono-trifluoromethyl-substituted ligand **1g** with Au(OAc)₃ gave a *ca.* 2 : 1 mixture of the two regioisomers **2g-p-Au(OAc^F)₂** and **2g-o-Au(OAc^F)₂** (Scheme 5). Thus, cyclometalation of Au(III) is compatible with the steric demands of an *ortho*-trifluoromethyl group. This result therefore suggests that the electronic influence is the main reason for the failure to produce **2f-Au(OAc^F)₂**, as the di(trifluoromethyl)-substituted ligand **1f** is more electron deficient than the corresponding mono-trifluoromethyl-substituted ligand **1g** ($\Delta C2' = +0.025$ for **1g-p-H⁺** and $\Delta C2' = +0.038$ for **1g-o-H⁺**).



Scheme 4 Synthesis of **1f-Au(OAc^F)₃**. The corresponding (N,C) complex **2f-Au(OAc^F)₂** was not observed.



Scheme 5 Reactivity of ligand **1g** towards Au(OAc)₃.

Synthesis of (N,C,C)-cyclometalated Au(III) trifluoroacetate pincer complexes

Following the successful microwave-mediated synthesis of pincer complex **3b-AuOAc^F** *via* C(sp²)-H and C(sp³)-H bond activation of ligand **1b** (Fig. 3),⁸ we sought to investigate related tridentate ligands (**1s-1u**) in order to get more insight into the scope and limitations of the pincer formation. DFT calculations on the formation of **3b-AuOAc^F** suggested the directing effect of the bulky *tert*-butyl group in ligand **1b** to be a key element in the successful synthesis of the complex.⁸ Therefore, ligands **1s** and **1t** with less sterically demanding substituents, but otherwise an identical substitution pattern to that of **1b** were investigated as tridentate ligands for Au(III). Additionally, the mono-substituted analogue of **1t**, ligand **1u**, was investigated in order to probe the selectivity of pincer formation relative to (N,C) cyclometalation.

Attempts of synthesizing **3s-AuOAc^F** in an analogous manner (microwave heating at 120 °C for 30 min) to **3b-AuOAc^F** failed, and the *N*-coordinated adduct of ligand **1s**, **1s-Au(OAc^F)₃**, was the main species observed in the ¹H NMR spectrum of the crude product. The combination of lower reaction temperature (80 °C) and longer reaction time (3.5 h) did however furnish the pincer complex **3s-AuOAc^F** in moderate yields (38%) (Scheme 6). It is noteworthy that the C(sp³)-H bond activation of the isopropyl group introduces a chiral

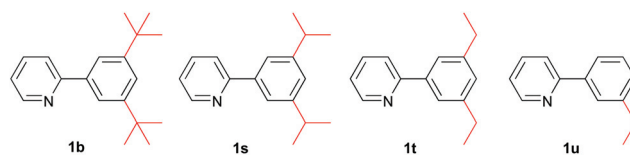
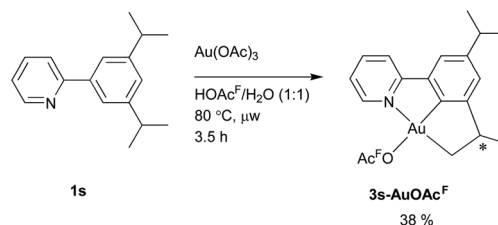


Fig. 3 Tridentate (N,C,C) ligands.



Scheme 6 Synthesis of **3s-AuOAc^F**. The star indicates the chiral centre that is formed upon C(sp³)-H bond activation.



centre in close proximity to gold in **3s-AuOAc^F**, which has previously been observed for a structurally related Pt(IV) complex.⁷⁹ We did not attempt to resolve the enantiomers, but the accessibility to chiral centres through cyclometalation is a topic that deserves further investigation, especially if an enantiopure complex can be obtained.⁸⁰

By employing the reaction conditions in Scheme 6, the ethyl analogue **3t-AuOAc^F** was obtained in a good yield (65%) from ligand **1t**. In addition to the synthesis of **3s-AuOAc^F** and **3t-AuOAc^F**, mono-ethyl-substituted **1u** was explored as a tridentate ligand for Au(III). The reaction of **1u** with Au(OAc)₃ yielded the desired pincer complex **3u-AuOAc^F** as a mixture with the corresponding (N,C)-cyclometalated complex **2u-p-Au(OAc^F)₂** (Scheme 7). The other possible (N,C)-cyclometalated complex **2u-o-Au(OAc^F)₂** could not be observed, supporting the involvement of **2u-o-Au(OAc^F)₂** as an intermediate for the formation of **3u-AuOAc^F** (see ESI† for details). To summarize, the successful syntheses of **3s-AuOAc^F** and **3t-AuOAc^F** show that the sterically induced pre-orientation of the C–H bond that is activated (as seen in ligand **1b**) is not a strict requirement for C(sp³)–H activation at Au(III), although the experimental observations suggest that the process is more feasible for **1b** compared to **1s** and **1t**. Importantly, the results show that cyclometalation of Au(III) through C(sp³)–H bond activation takes place at a β position regardless of the nature of the alkyl group in the ligand.

Single-crystal X-ray diffraction analysis of (N,C)- and (N,C,C)-cyclometalated Au(III) complexes

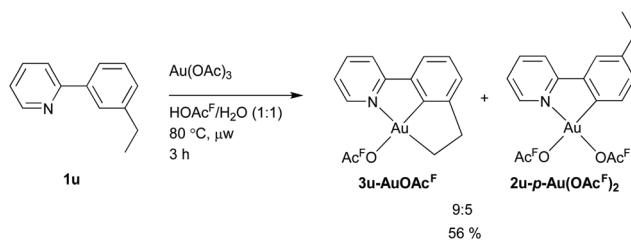
Several of the (N,C)-cyclometalated Au(III) complexes were analysed by single-crystal X-ray diffraction (Fig. 4). In all cases, the complexes crystallized with the expected square planar geometry around Au(III). Deviations from ideal square planar geometries were evaluated by the τ₄ values (0 for square planar complexes, 1 for tetrahedral complexes)^{81,82} of the structures, which were found in the range from 0.02 to 0.09 (Table 1). All Au–ligand bond distances and angles are given in Table 1. The bond angles are similar to those previously reported for structurally related square planar Au(III) complexes.^{19,21,26,29} The Au1–N1 and Au1–C7 bond lengths are in the range of 1.9974(19)–2.0426(17) Å and 1.988(2)–2.02(2) Å, being comparable to those reported for **2a-Au(OAc^F)₂** (1.991(6) Å and 1.995(7) Å, respectively).¹⁹ The Au1–O3 (*trans* to N) bond lengths are in the range of 2.0032(16)–2.034(3) Å, all being slightly longer than the corresponding bond length in **2a-Au(OAc^F)₂** (1.993(5)

Å). Larger variation is found in the Au1–O1 (*trans* to C) bond lengths. In **2c-Au(OAc^F)₂**, the Au1–N1 and Au1–C7 bond lengths are comparable to the other complexes, but a slightly shorter Au1–O1 (*trans* to carbon) bond is noted (2.070(3) Å for **2c-Au(OAc^F)₂** vs. 2.095(13)–2.1185(16) Å for the rest of the complexes). This may be attributed to the presence of the electron-withdrawing fluorine substituents in the ligand, making the phenyl ring a weaker *trans* influence ligand compared to the one in e.g. **2a-Au(OAc^F)₂**. Similar, but less pronounced, shortening of the Au1–O1 bond can be observed for other complexes having electron-withdrawing substituents (**2e-Au(OAc^F)₂**, **2o-Au(OAc^F)₂** and **2r-Au(OAc^F)₂**; Au1–O1 = 2.1026(18) Å, 2.0984(16) Å and 2.095(13) Å, respectively).

For 3',5'-disubstituted complexes **2c-Au(OAc^F)₂**, **2h-Au(OAc^F)₂** and **2r-Au(OAc^F)₂** the O3–Au1–C7 *cis* angles (98.89(14)°, 99.68(8)° and 98.0(6)°, respectively) were slightly larger than in the other complexes (see Table 1). The angle was larger for **2h-Au(OAc^F)₂** than for **2r-Au(OAc^F)₂**, being in accordance with literature reports of the similar effective steric bulk of the pentafluorophenyl group and the methyl group.⁸³

The Au1–N1 bond lengths of **2i-Au(OAc^F)₂** and **2m-Au(OAc^F)₂** are 2.034(3) Å and 2.0426(17) Å, respectively, being longer than the corresponding Au1–N1 bond in the other complexes studied. Furthermore, the N1–Au1–O1 *cis* angles in **2i-Au(OAc^F)₂** and **2m-Au(OAc^F)₂** are also larger (103.80(11)° and 104.62(6)°, respectively) compared to the other complexes (91.37(12)°–99.00(8)°). Similar elongation of the Au1–N1 bond and widening of the corresponding *cis* angle have earlier been reported by Cinellu and co-workers for a structurally similar (N,C) Au(III) complex derived from 6,6'-dimethoxy-2,2'-bipyridine.⁸⁴ The N1–C1–O5 angle in **2m-Au(OAc^F)₂** is 114.51(18)°, being slightly smaller than the corresponding N–C–O(alkoxy) angles in reported crystal structures of non-coordinated pyridines, bipyridines or phenanthrolines with alkoxy substituents α to nitrogen (typically around 120°).^{40,84–88} The angle is similar to the N–C–O(alkoxy) angle in (N,C) Au(III) and Pd(II) complexes derived from 6,6'-dimethoxy-2,2'-bipyridine reported by Cinellu and co-workers.^{40,84} The relatively small N–C–O(alkoxy) angle in **2m-Au(OAc^F)₂**, together with a Au1–O5 distance of 3.13 Å may hint at a weak interaction between the gold centre and the oxygen atom in the methoxy group. This was further investigated by means of ¹⁵N NMR spectroscopy (see below). The N1–C1–C13 angle (120.9(3)°) in 6-methyl-substituted **2i-Au(OAc^F)₂** is slightly larger than the corresponding N–C–C(alkyl) angles in reported crystal structures of non-coordinated pyridines, bipyridines or phenanthrolines with alkyl substituents α to nitrogen (typically around 115°).^{89–93} This indicates that Au(III) coordination to 6-methyl-substituted pyridines results in a widening of the N–C–C(methyl) angle, contrary to the effect Au(III) coordination has on the N–C–O(methoxy) angle in 6-methoxy-substituted pyridines. The N–C–C(alkyl) angle in **2i-Au(OAc^F)₂** is similar to reported angles for related square planar Au(III)⁹⁴ and Pd(II)⁵⁸ complexes of 6-methyl-substituted (bi)pyridine ligands.

Au(III) pincer complexes **3s-AuOAc^F** and **3t-AuOAc^F** were also crystallographically characterized (Fig. 5). Both complexes crys-



Scheme 7 Reactivity of ligand **1u** towards Au(OAc)₃.



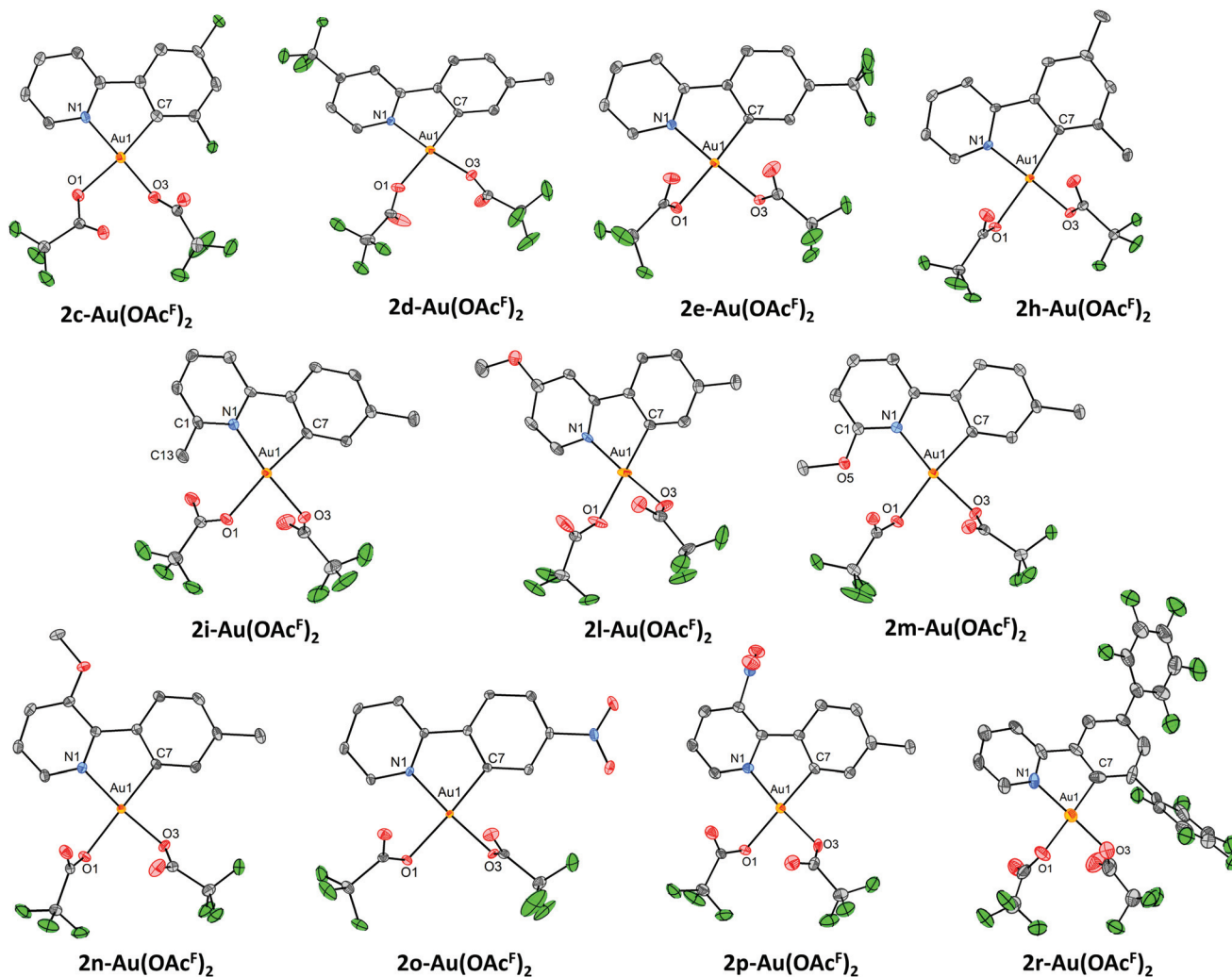


Fig. 4 ORTEPs of $2c\text{-Au}(\text{OAc}^{\text{F}})_2$, $2d\text{-Au}(\text{OAc}^{\text{F}})_2$, $2e\text{-Au}(\text{OAc}^{\text{F}})_2$, $2h\text{-Au}(\text{OAc}^{\text{F}})_2$, $2i\text{-Au}(\text{OAc}^{\text{F}})_2$, $2l\text{-Au}(\text{OAc}^{\text{F}})_2$, $2m\text{-Au}(\text{OAc}^{\text{F}})_2$, $2n\text{-Au}(\text{OAc}^{\text{F}})_2$, $2o\text{-Au}(\text{OAc}^{\text{F}})_2$, $2p\text{-Au}(\text{OAc}^{\text{F}})_2$ and $2r\text{-Au}(\text{OAc}^{\text{F}})_2$. Ellipsoids are shown at 50% probability level. Hydrogen atoms and associated (disordered) solvent molecules (for $2d\text{-Au}(\text{OAc}^{\text{F}})_2$, $2i\text{-Au}(\text{OAc}^{\text{F}})_2$, $2m\text{-Au}(\text{OAc}^{\text{F}})_2$ and $2p\text{-Au}(\text{OAc}^{\text{F}})_2$) have been removed for clarity. For $2d\text{-Au}(\text{OAc}^{\text{F}})_2$ and $2i\text{-Au}(\text{OAc}^{\text{F}})_2$, only one of the two molecules in the asymmetric unit of each structure is shown. For $2d\text{-Au}(\text{OAc}^{\text{F}})_2$ and $2l\text{-Au}(\text{OAc}^{\text{F}})_2$, disorder in the trifluoroacetate ligands have been removed for clarity. For more details, see ESI.†

Table 1 τ_4 values and selected bond lengths (Å) and angles ($^\circ$) for (N,C) Au(III) complexes

Complex	τ_4	Au1–N1	Au1–C7	Au1–O1	Au1–O3	N1–Au1–C7	N1–Au1–O1	N1–Au1–O3	O1–Au1–O3	O1–Au1–C7	O3–Au1–C7
$2a\text{-Au}(\text{OAc}^{\text{F}})_2$	0.07	1.991(6)	1.995(7)	2.111(5)	1.993(5)	81.8(3)	93.1(2)	175.5(2)	88.8(2)	174.8(3)	96.4(3)
$2c\text{-Au}(\text{OAc}^{\text{F}})_2$	0.08	2.003(3)	2.004(4)	2.070(3)	1.999(3)	81.86(14)	91.37(12)	175.45(12)	87.60(11)	172.50(13)	98.89(14)
$2d\text{-Au}(\text{OAc}^{\text{F}})_2$	0.06	2.013(4)	1.995(5)	2.114(4)	2.008(4)	82.03(19)	94.74(17)	174.92(17)	88.41(17)	176.60(19)	94.73(18)
$2e\text{-Au}(\text{OAc}^{\text{F}})_2$	0.04	2.007(2)	1.991(2)	2.1026(18)	2.0065(18)	81.83(9)	95.48(8)	175.98(8)	87.67(7)	177.24(9)	95.05(9)
$2h\text{-Au}(\text{OAc}^{\text{F}})_2$	0.04	1.9996(19)	2.022(2)	2.1185(16)	2.0085(16)	81.84(8)	93.69(7)	178.23(7)	84.76(7)	175.38(8)	99.68(8)
$2i\text{-Au}(\text{OAc}^{\text{F}})_2$	0.09	2.034(3)	1.994(3)	2.113(2)	2.011(2)	81.65(13)	103.80(11)	173.84(11)	82.24(10)	172.37(12)	92.46(12)
$2l\text{-Au}(\text{OAc}^{\text{F}})_2$	0.05	1.999(4)	2.000(5)	2.101(4)	2.005(4)	81.77(17)	96.02(17)	176.75(16)	86.37(18)	176.63(17)	95.74(18)
$2m\text{-Au}(\text{OAc}^{\text{F}})_2$	0.09	2.0426(17)	1.995(2)	2.1112(15)	2.0119(15)	81.23(8)	104.62(6)	171.95(6)	83.31(6)	173.98(7)	90.81(7)
$2n\text{-Au}(\text{OAc}^{\text{F}})_2$	0.03	1.9974(19)	1.994(2)	2.1183(17)	2.0032(16)	81.43(9)	99.00(8)	173.40(7)	86.39(7)	179.55(8)	93.19(8)
$2o\text{-Au}(\text{OAc}^{\text{F}})_2$	0.06	2.0047(18)	1.988(2)	2.0984(16)	2.0097(16)	81.62(8)	95.05(7)	175.35(7)	89.37(7)	176.23(7)	94.01(8)
$2p\text{-Au}(\text{OAc}^{\text{F}})_2$	0.03	2.008(3)	1.996(3)	2.102(2)	2.034(3)	81.54(12)	98.36(10)	173.63(11)	86.78(9)	179.69(11)	93.29(12)
$2r\text{-Au}(\text{OAc}^{\text{F}})_2$	0.02	1.999(12)	2.020(19)	2.095(13)	1.984(11)	82.3(7)	95.7(5)	178.9(5)	84.0(5)	177.0(6)	98.0(6)

The corresponding data for $2a\text{-Au}(\text{OAc}^{\text{F}})_2$ are included for reference purposes.¹⁹ For $2d\text{-Au}(\text{OAc}^{\text{F}})_2$ and $2i\text{-Au}(\text{OAc}^{\text{F}})_2$, metric data for one of the two molecules in the respective asymmetric unit are listed. See ESI (Fig. S140 and Fig. S145)† for metric data for both molecules in each asymmetric unit.



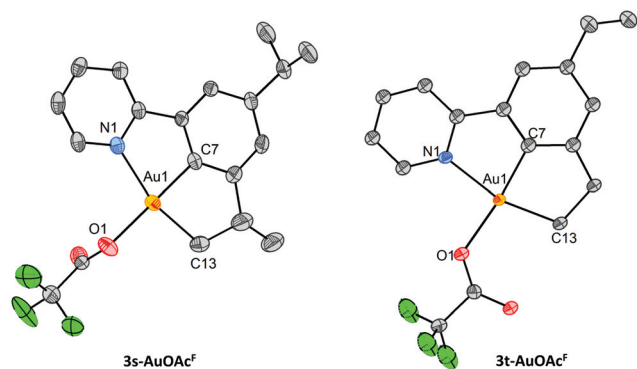


Fig. 5 ORTEPs of **3s-AuOAc^F** and **3t-AuOAc^F**. Ellipsoids are shown at 50% probability level. Hydrogen atoms and disorder in the isopropyl group in **3s-AuOAc^F** have been omitted for clarity. Only one of the two molecules in the asymmetric unit is shown for **3s-AuOAc^F**, and only one of the four molecules in the asymmetric unit is shown for **3t-AuOAc^F**. For more details, see ESI†

tallized with distorted square planar geometry around Au(III) ($\tau'_4 = 0.09$ for **3s-AuOAc^F** and $\tau'_4 = 0.11$ for **3t-AuOAc^F**). The bond lengths and angles of the complexes are similar to those reported for **3b-AuOAc^F** (Table 2). Interestingly, only one enantiomer of the racemic mixture of pincer complex **3s-AuOAc^F** could be modelled as the major component during refinement of the single-crystal X-ray structure.

¹⁵N NMR spectroscopic studies of cyclometalated Au(III) complexes

Some of the complexes and ligands discussed herein were investigated by ¹⁵N NMR spectroscopy, and coordination shifts $\Delta\delta^{15}\text{N}$ ($\delta^{15}\text{N}_{\text{complex}} - \delta^{15}\text{N}_{\text{ligand}}$) were obtained in order to gain insight about the Au–N interactions. Furthermore, a selection of previously reported tpy-ligated Au(III) complexes^{8,14,19,22,25,95} with varying substituents *cis* and *trans* to pyridine-N was studied by ¹⁵N NMR spectroscopy to shed light on which factors influence the ¹⁵N NMR chemical shifts of *N*-ligated square planar d⁸ metal complexes (Fig. 6).

$\Delta\delta^{15}\text{N}$ were found in the range of –88.2 ppm to –104.5 ppm for the (N,C) di(trifluoroacetate) Au(III) complexes (except for **2m-Au(OAc^F)₂**; see discussion below). These shifts are similar to reported data for other pyridine-ligated Au(III) complexes with weak *trans* influence ligands *trans* to nitrogen.^{44,96–98} The ¹⁵N NMR data can be interpreted in a

similar manner as the data from the single-crystal X-ray diffraction analysis of the complexes, where the Au–N bond lengths were found to be little dependent of the substituents in the ligand backbone. This reflects that other factors, such as the ligand *trans* to pyridine-N,⁹⁹ affects $\Delta\delta^{15}\text{N}$ stronger than the substituents on the pyridine ring. Similar observations have also been reported by Pazderski for square planar Pd(II) and Pt(II) complexes with substituted bipyridine and phenanthroline ligands.⁹⁶

The coordination shift of the 6-methoxy-substituted complex **2m-Au(OAc^F)₂** ($\Delta\delta^{15}\text{N} = -66.1$ ppm) is significantly smaller than those of the other substituted di(trifluoroacetate) complexes. It seems likely that the relatively small $\Delta\delta^{15}\text{N}$ found for **2m-Au(OAc^F)₂** is a result of the weak interaction between the methoxy-oxygen and gold. This potential interaction was also observed in the single-crystal structure of the Au(III) complex. The coordination shift for the 6-methyl-substituted complex **2i-Au(OAc^F)₂** (–88.2 ppm) was in the same range as those obtained for the other di(trifluoroacetate) complexes studied herein, although the Au–N bond length in the crystal structure of **2i-Au(OAc^F)₂** (2.034(3) Å) was similar to the one observed for **2m-Au(OAc^F)₂** (2.0426(17) Å). The very similar Au–N bond lengths in the two complexes strengthens the argument that the relatively small coordination shift obtained for **2m-Au(OAc^F)₂** is caused by an interaction between oxygen and gold, rather than steric repulsion between gold and the 6-substituent.⁸⁴ If this was the case, it would be expected to yield a comparable coordination shift for **2i-Au(OAc^F)₂**.

For complex **1f-Au(OAc^F)₃**, $\Delta\delta^{15}\text{N}$ was found to be –114.1 ppm, larger than what was observed for the cyclometalated complexes and also larger than what has been reported for pyridine-ligated AuCl₃ complexes in the literature (*ca.* –80 ppm in CDCl₃).^{100–102} This reflects that [OAc^F][–] is a weaker *trans* influence ligand than Cl[–], being consistent with reported experimental and computational data for the *trans* influence of carboxylate ligands *vs.* chloride ligands in square planar complexes.^{103–105}

To further evaluate the effect of the identity of the ligand *trans* to nitrogen on the coordination shift of pyridine-N, $\Delta\delta^{15}\text{N}$ for **2a-Au(OAc)₂** and **2a-Au(CH₃)₂** were obtained, and compared to the one found for **2a-Au(OAc^F)₂**. For the three complexes, $\Delta\delta^{15}\text{N}$ was found to decrease in the order **2a-Au(OAc^F)₂** \sim **2a-Au(OAc)₂** > **2a-Au(CH₃)₂**, agreeing with the established *trans* influence of the corresponding ligands; [OAc^F][–] \sim [OAc][–] < CH₃[–].^{19,103,104} Coordination shifts were also obtained

Table 2 τ'_4 values and selected bond lengths (Å) and angles (°) for (N,C,C) Au(III) complexes

Complex	τ'_4	Au1–N1	Au1–C7	Au1–O1	Au1–C13	N1–Au1–C7	N1–Au1–C13	N1–Au1–O1	C7–Au1–C13	C7–Au1–O1	C13–Au1–O1
3b-AuOAc^F	0.08	2.135(3)	1.944(3)	2.119(3)	2.049(4)	80.36(13)	161.36(14)	99.62(11)	81.66(15)	179.34(13)	98.30(14)
3s-AuOAc^F	0.09	2.144(3)	1.954(3)	2.108(3)	2.055(4)	80.59(13)	161.67(14)	98.52(11)	81.33(15)	177.53(13)	99.67(14)
3t-AuOAc^F	0.11	2.15(2)	1.94(2)	2.076(18)	2.04(3)	80.0(10)	163.0(9)	93.2(8)	83.0(11)	173.2(10)	103.7(9)

The corresponding data for **3b-AuOAc^F** are included for reference purposes.⁸ For **3s-AuOAc^F**, metric data for one of the two molecules in the asymmetric unit are listed. See Fig. S160 (ESI)† for metric data for both molecules in the asymmetric unit. For **3t-AuOAc^F**, metric data for one of the four molecules in the asymmetric unit are listed. See Fig. S162 (ESI)† for metric data for all molecules in the asymmetric unit.



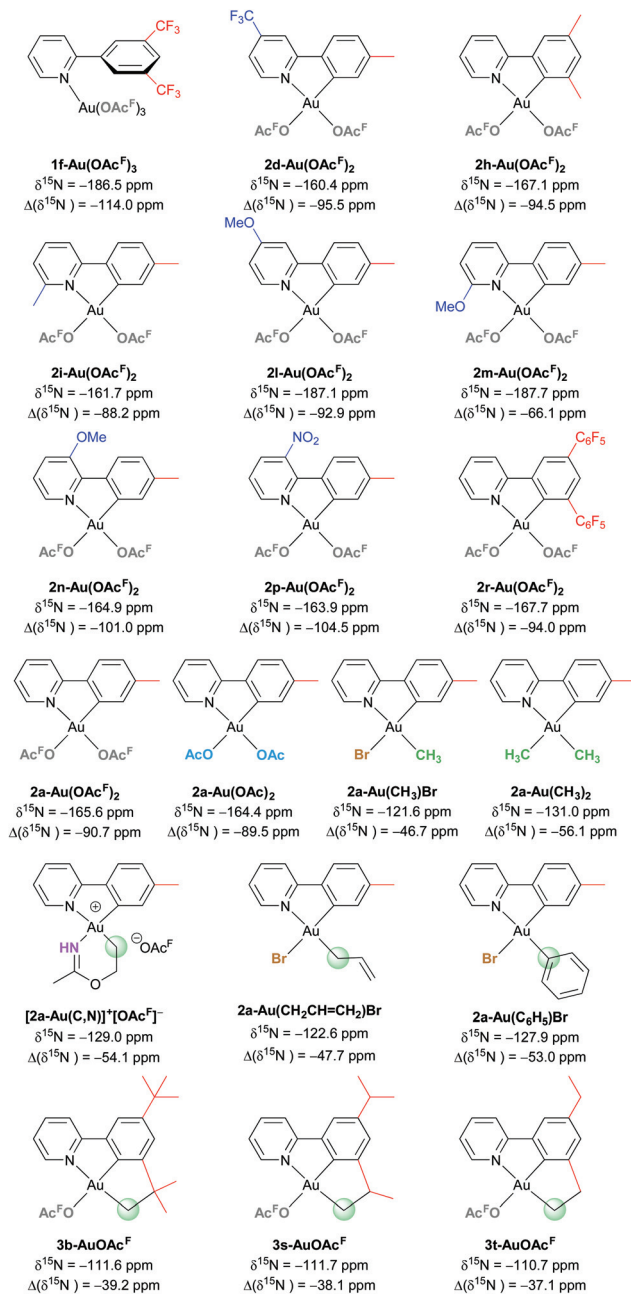


Fig. 6 Overview of (N,C) and (N,C,C) Au(III) complexes studied by ^{15}N NMR spectroscopy herein. All data were collected in CD_2Cl_2 at either 600 or 800 MHz. See also Table S1, ESI[†]. $\delta^{15}\text{N}$ for $[\text{2a-Au(C,N)}]^+[\text{OAcF}]^-$ is from ref. 25.

for the pincer complexes $3\text{b-AuOAc}^{\text{F}}$, $3\text{s-AuOAc}^{\text{F}}$ and $3\text{t-AuOAc}^{\text{F}}$ which all have an alkyl group *trans* to pyridine-N. As expected from the differences in relative *trans* influence strength of an alkyl ligand and a $[\text{OAcF}]^-$ ligand, the coordination shifts for the three pincers were significantly smaller ($\Delta\delta^{15}\text{N}$ from -37.1 to -39.5 ppm) than those obtained for the di(trifluoroacetate) complexes. Surprisingly, they were also found to be smaller than the coordination shifts of $2\text{a-Au(CH}_3)_2$, $2\text{a-Au(CH}_2\text{CH=CH}_2)\text{Br}$, $2\text{a-Au(CH}_3)\text{Br}$ and $[\text{2a-Au(C,N)}]^+[\text{OAcF}]^-$,

although it could be anticipated that the relative *trans* influence of the ligands *trans* to pyridine-N would be similar for all these complexes. The observations may be explained from differences in the relative *cis* influence of an alkyl ligand, a halide ligand and a carboxylate ligand. A smaller $\Delta\delta^{15}\text{N}$ was found for $2\text{a-Au(CH}_3)\text{Br}$ ($\Delta\delta^{15}\text{N} = -46.7$ ppm) compared to $2\text{a-Au(CH}_3)_2$ ($\Delta\delta^{15}\text{N} = -56.1$ ppm) being in accordance with the reported higher *cis* influence of halide ligands compared to alkyl ligands.^{106,107} In summary, the findings from the ^{15}N NMR spectroscopic studies show that functionalization of the ligand backbone has little effect on the interaction between pyridine-N and gold, whereas the nature of the ligands *cis* and *trans* to pyridine-N has a significantly larger effect.

DFT calculations on the formation of (N,C,C)-cyclometalated Au(III) complexes $3\text{s-AuOAc}^{\text{F}}$ and $3\text{t-AuOAc}^{\text{F}}$

In order to gain understanding of pincer formation for tridentate (N,C,C) ligands **1s** and **1t**, DFT calculations were performed. The formation of $3\text{s-AuOAc}^{\text{F}}$ and $3\text{t-AuOAc}^{\text{F}}$ starting from complexes 2s-Au(OAcF)_2 and 2t-Au(OAcF)_2 via the same mechanism proposed for the formation of $3\text{b-AuOAc}^{\text{F}}$ ⁸ was explored (Fig. 7). As the formation of 2s-Au(OAcF)_2 was found to be very similar to that of 2b-Au(OAcF)_2 , the energies for this first cyclometalation step are not included in the figure (see Table S17, ESI[†] for details), and was not calculated for 2t-Au(OAcF)_2 .

Looking at the C–H activation step, a clear difference in energy for TS1 was observed for the three complexes, illustrating the effect of the substituent on the formation of the desired pincer complex. The endergonic dissociation step forming the agostic intermediates $[\text{2s-Au-OAc}^{\text{F}}]^+[\text{OAcF}]^-$ and $[\text{2t-Au-OAc}^{\text{F}}]^+[\text{OAcF}]^-$ is more than 10 kcal mol⁻¹ higher in

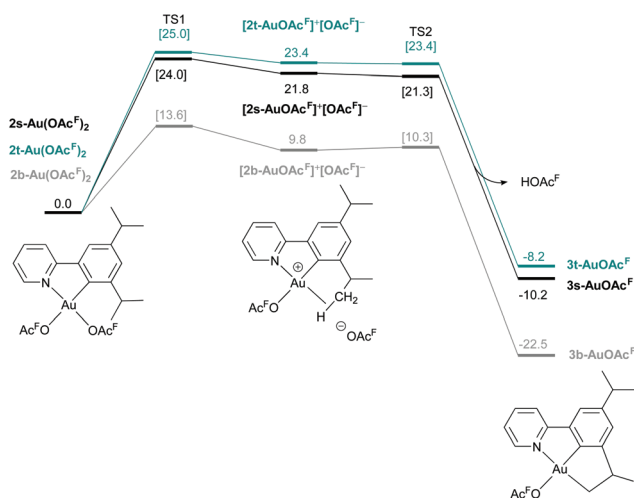
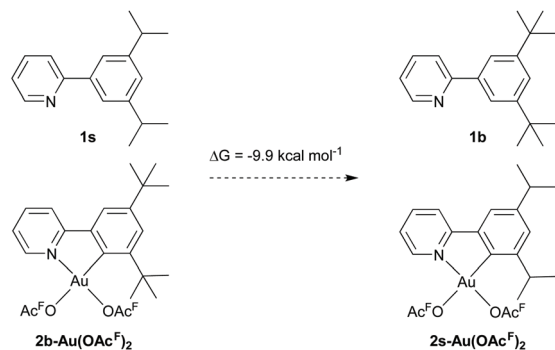


Fig. 7 Free energy profile in kcal mol⁻¹ for the formation of $3\text{b-AuOAc}^{\text{F}}$, $3\text{s-AuOAc}^{\text{F}}$ and $3\text{t-AuOAc}^{\text{F}}$ from the corresponding di(trifluoroacetate) complexes. In the figure, 2s-Au(OAcF)_2 , $[\text{2s-AuOAc}^{\text{F}}]^+[\text{OAcF}]^-$ and $3\text{s-AuOAc}^{\text{F}}$ are displayed as structural examples. The energies of all minima and transition states in brackets are computed in CH_2Cl_2 (SMD) for $3\text{s-AuOAc}^{\text{F}}$ and $3\text{t-AuOAc}^{\text{F}}$. The energies and transition states for $3\text{b-AuOAc}^{\text{F}}$ were computed in HOAc^{F} .





Scheme 8 Isodesmic reaction for the formal chelate ligand exchange on $2b\text{-Au}(\text{OAc}^{\text{F}})_2$ furnishing $2s\text{-Au}(\text{OAc}^{\text{F}})_2$. Gibbs energy in CH_2Cl_2 (SMD) is given in kcal mol^{-1} .

energy for $3s\text{-AuOAc}^{\text{F}}$ and $3t\text{-AuOAc}^{\text{F}}$ ($24.0 \text{ kcal mol}^{-1}$ and $25.0 \text{ kcal mol}^{-1}$, respectively), compared to $3b\text{-AuOAc}^{\text{F}}$ ($13.6 \text{ kcal mol}^{-1}$). The following proton abstractions by $[\text{OAc}^{\text{F}}]^-$ that furnish $3s\text{-AuOAc}^{\text{F}}$ and $3t\text{-AuOAc}^{\text{F}}$ are barrier-free, which is in accordance with what was found for $3b\text{-AuOAc}^{\text{F}}$.

The energy barrier associated with the formation of the agostic intermediates is depending on the bulkiness of the alkyl substituent that undergoes C–H activation. The lower energy barrier for $[2b\text{-Au-OAc}^{\text{F}}]^+[\text{OAc}^{\text{F}}]^-$ is due to the higher energy of $2b\text{-Au}(\text{OAc}^{\text{F}})_2$ relative to TS1 due to steric interaction with the *tert*-butyl group in the (N,C)-cyclometalated complex. In order to highlight the difference in stability of the (N,C)-cyclometalated complexes $2b\text{-Au}(\text{OAc}^{\text{F}})_2$ and e.g. $2s\text{-Au}(\text{OAc}^{\text{F}})_2$, which is responsible for the significantly lower TS1 found for the *tert*-butyl system, an isodesmic reaction of a formal chelate ligand exchange on $2b\text{-Au}(\text{OAc}^{\text{F}})_2$ with **1s** was investigated (Scheme 8). The formation of $2s\text{-Au}(\text{OAc}^{\text{F}})_2$ and **1b** is favoured by $9.9 \text{ kcal mol}^{-1}$ which illustrates the negative effect the large, bulky substituents have on the stability of di(trifluoroacetate) complexes. On the other hand, this characteristic of the *tert*-butyl substituent ultimately facilitates the $\text{C}(\text{sp}^3)\text{-H}$ bond activation step and subsequent pincer formation.

Conclusions

In this work, a detailed experimental and computational study of a series of 2-arylpyridine-based (N,C)- and (N,C,C)-cyclometalated Au(III) complexes has been presented. For the (N,C) systems, it was found that the scope of microwave-mediated synthesis of cyclometalated Au(III) complexes is broad and that a large variety of different functional groups is tolerated. This makes it an attractive method for the synthesis of Au(III) complexes of substituted arylpyridine ligands without having to resort to any organomercury intermediates. The efficiency of the reaction is strongly dependent on the electronic features of the (N,C) ligand, being consistent with an electrophilic aromatic substitution-type mechanism for cyclometalation at Au(III). Natural charge analysis performed on the protonated (N,C) ligands was found to correlate with the experimental obser-

vations of their reactivity towards $\text{Au}(\text{OAc})_3$. Single-crystal X-ray diffraction and ^{15}N NMR spectroscopy data suggest that the (N,C) Au(III) complexes are structurally similar species, meaning that the ligand scaffold is flexible to changes without drastically affecting the coordination sphere around Au. Detailed ^{15}N NMR spectroscopic studies of different cyclometalated Au(III) complexes with varying ligands *trans* to the (N,C) backbone show that these ligands have a much stronger influence on the Au–N interaction than any substituents in the backbone, with the exception of the 6-methoxy-substituted complex $2m\text{-Au}(\text{OAc}^{\text{F}})_2$. For this complex, the relatively small coordination shift may be explained by a possible weak interaction between the methoxy-oxygen and gold. In addition to the studies of (N,C)-cyclometalated Au(III) complexes, $\text{C}(\text{sp}^3)\text{-H}$ bond activation as a synthetic feasible method to yield (N,C,C)-cyclometalated Au(III) complexes was expanded. Earlier we have shown that the ligand 2-(3,5-di-*tert*-butylphenyl)pyridine (**1b**) functions as a tridentate ligand for Au(III), yielding pincer complex $3b\text{-AuOAc}^{\text{F}}$. We have broadened the scope of Au(III) pincer formation, *via* $\text{C}(\text{sp}^3)\text{-H}$ bond activation to include less sterically encumbered ligands. These ligands contain either isopropyl (**1s**) or ethyl (**1t** and **1u**) groups, where the C–H bond that is activated is able to rotate away from gold. In a combined experimental and computational effort we have shown that these ligands indeed undergo $\text{C}(\text{sp}^3)\text{-H}$ bond activation, but that the process is less facile than for the *tert*-butyl-substituted system. Further work will focus on broadening the scope of (N,C,C) pincer formation from alkyl-substituted 2-arylpyridines, as well as investigating their reactivity and optical properties.

Experimental section

General considerations

$2a\text{-Au}(\text{OAc}^{\text{F}})_2$,¹⁹ $3b\text{-AuOAc}^{\text{F}}$,⁸ $2a\text{-Au}(\text{OAc})_2$,^{14,95} $2a\text{-Au}(\text{CH}_2\text{CH}=\text{CH}_2)\text{Br}$,²² $2a\text{-Au}(\text{C}_6\text{H}_5)\text{Br}$,¹⁹ $2a\text{-Au}(\text{CH}_3)\text{Br}$,¹⁹ $2a\text{-Au}(\text{CH}_3)_2$,¹⁹ 3,5-diethylphenylboronic acid^{108,109} and 3,5-diisopropylphenylboronic acid^{108,109} were synthesized according to literature procedures. $\text{Au}(\text{OAc})_3$ was obtained from aber. THF (unstabilized) and CH_2Cl_2 were dried using an MB SPS-800 solvent purifier system from MBraun. Hexanes and ethyl acetate were distilled before use. Deionized water was used. Other chemicals and solvents were used as received from commercial sources. TLC was performed using Merck 60 F254 plates. Flash chromatography was performed using silica gel from Merck (60, 0.040–0.063 mm). Microwave reactions were performed with a Milestone MicroSYNTH microwave reactor with a SK-10 rotor or, for reaction volumes smaller than 10 mL, in an Anton Paar GmbH Monowave 300 synthesis reactor equipped with an internal IR probe calibrated with a Ruby thermometer. NMR spectroscopy was performed using Bruker Avance DPX300, AVII400, AVIIIHD400, DRX500, AVI600, AVII600 or AVIIIHD800 operating at 300 MHz (^1H NMR), or 400 MHz (^1H NMR), 376 MHz (^{19}F NMR), 101 MHz (^{13}C NMR), or 500 MHz (^1H NMR), or 600 MHz (^1H NMR) and 151 MHz



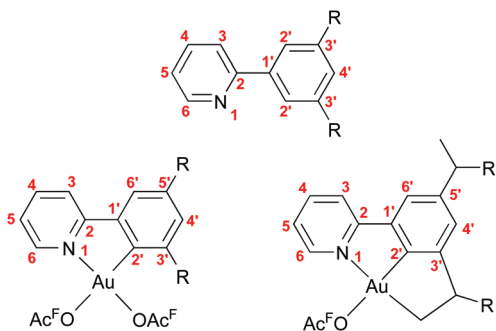


Fig. 8 Numbering scheme used for reporting the NMR data.

(^{13}C NMR), or 800 MHz (^1H NMR) and 201 MHz (^{13}C NMR) respectively. All spectra were recorded at room temperature. ^1H NMR and ^{13}C NMR spectra have been referenced relative to the residual solvent signals, and the resonances are numbered according to Fig. 8. Chemical shifts in ^{19}F NMR have been referenced to CFCl_3 by using C_6F_6 or $\text{C}_6\text{H}_5\text{F}$ (-164.9 ppm and -116.1 ppm with respect to CFCl_3 at 0 ppm) as an internal standard, and are proton decoupled. Chemical shifts in ^{15}N NMR have been calibrated against CH_3NO_2 as an external standard (0.0 ppm). All ^{15}N NMR chemical shifts were obtained and assigned using ^1H - ^{15}N HMBC experiments. The resonances in the ^1H NMR and ^{13}C NMR spectra were assigned using various 2D experiments (COSY, NOESY, HSQC and HMBC). MS (ESI and APPI) was recorded on a Bruker maXis II ETD spectrometer. All melting points are uncorrected and were obtained with a Stuart SMP10 melting point apparatus. Elemental analysis was performed by Mikroanalytisches Laboratorium Kolbe, Oberhausen, Germany. Single-crystal diffraction data were acquired on a Bruker D8 Venture equipped with a Photon 100 CMOS area detector, and using Mo $\text{K}\alpha$ radiation ($\lambda = 0.71073$ Å) or Cu $\text{K}\alpha$ radiation ($\lambda = 1.5406$ Å) from an Incoatec μS microsource. Data reduction was performed with the Bruker Apex3 Suite, the structures were solved with ShelXT¹¹⁰ and refined with ShelXL.¹¹¹ Olex2 was used as user interface.¹¹² The cif files were edited with enCIFer v. 1.4.¹¹³ Full details of the data collection, structure solution and refinement for each compound are contained in the cif files, available from [Experimental and analytical data for a selection of compounds described within the text are presented here, data for all compounds can be found in the ESI.†](https://www.ccdc.cam.ac.uk/(CCDC 2115512 (2c-Au(OAc^F)₂), 2122527 (2d-Au(OAc^F)₂), 2085151 (2e-Au(OAc^F)₂), 2086346 (2h-Au(OAc^F)₂), 2126159 (2i-Au(OAc^F)₂), 2126283 (2j₃Au₂(OAc^F)₂), 2122284 (2l-Au(OAc^F)₂), 2111917 (2m-Au(OAc^F)₂), 2105655 (2n-Au(OAc^F)₂), 2086931 (2o-Au(OAc^F)₂), 2126114 (2p-Au(OAc^F)₂), 2130186 (2r-Au(OAc^F)₂), 2126097 (3s-Au(OAc^F)₂) and 2114274 (3t-Au(OAc^F)₂)). The data are summarized in Tables S2–S15, ESI.†</p>
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General procedure for synthesis of arylpyridine ligands

2-Bromopyridine or substituted 2-bromopyridine (5.00 mmol, 1.0 equiv.) and arylboronic acid (4.75–5.50 mmol, 0.95–1.1

equiv.) were dissolved in *n*-PrOH (10 mL). A solution of K_3PO_4 (10.0–11.5 mmol, 2.0–2.3 equiv.) in water (10 mL) was added, and the resulting biphasic mixture was degassed for 10 min by bubbling Ar through it. $\text{Pd}(\text{OAc})_2$ (0.100 mmol, 2.0 mol%) and PPh_3 (0.300 mmol, 6.0 mol%) were added, and the reaction mixture was heated at reflux temperature for 3 h under Ar. After cooling to rt, CH_2Cl_2 (50 mL) and water (50 mL) were added. The phases were separated, and the CH_2Cl_2 solution was washed with 2 M NaOH (aq) (2×50 mL), brine (50 mL), and was dried over Na_2SO_4 . The solvent was removed under reduced pressure. The obtained residue was purified by flash column chromatography (hexanes/EtOAc, hexanes/ CH_2Cl_2 or hexanes/EtOAc/ CH_2Cl_2 mixtures), furnishing the arylpyridine ligand.

2-(3,5-Difluorophenyl)pyridine (1c). The general procedure was followed. 2-Bromopyridine (0.789 g, 4.99 mmol, 1.0 equiv.), 3,5-difluorophenylboronic acid (0.866 g, 5.48 mmol, 1.1 equiv.), K_3PO_4 (2.14 g, 10.1 mmol, 2.0 equiv.), $\text{Pd}(\text{OAc})_2$ (0.0236 g, 0.105 mmol, 2.1 mol%) and PPh_3 (0.0808 g, 0.308 mmol, 6.2 mol%) were used. The crude product was purified by flash column chromatography (98% hexanes/2% EtOAc as eluent to 95% hexanes/5% EtOAc) yielding **1c** as a colourless solid (0.870 g, 4.55 mmol, 91%). Mp. 61–62 °C. ^1H NMR (600 MHz, CDCl_3): δ 8.70 (d, $^3J_{\text{H,H}} = 4.8$ Hz, 1H, H⁶), 7.78 (ddd, $^3J_{\text{H,H}} = 7.8$ Hz, $^3J_{\text{H,H}} = 7.7$ Hz, $^4J_{\text{H,H}} = 1.8$ Hz, 1H, H⁴), 7.69 (d, $^3J_{\text{H,H}} = 7.9$ Hz, 1H, H³), 7.53–7.57 (m, 2H, H^{2'}), 7.29 (ddd, $^3J_{\text{H,H}} = 7.4$ Hz, $^3J_{\text{H,H}} = 4.8$ Hz, $^4J_{\text{H,H}} = 0.9$ Hz, 1H, H⁵), 6.85 ppm (m, 1H, H⁴). ^{13}C NMR (151 MHz, CDCl_3): δ 163.4 (dd, $^1J_{\text{C,F}} = 248.1$ Hz, $^3J_{\text{C,F}} = 12.4$ Hz, C³), 154.8 (t, $^4J_{\text{C,F}} = 2.8$ Hz, C²), 149.8 (C⁶), 142.7 (t, $^3J_{\text{C,F}} = 9.3$ Hz, C¹), 137.0 (C⁴), 123.2 (C⁵), 120.5 (C³), 109.6–109.8 (m, C^{2'}), 104.1 ppm (t, $^2J_{\text{C,F}} = 25.7$ Hz, C^{4'}). ^{19}F NMR (376 MHz, CDCl_3): δ -112.8 ppm (Ar-F). ^{15}N (^1H) NMR (600 MHz, CDCl_3): δ -72.2 ppm (N¹). MS (ESI): *m/z* (rel. %): 192.062 (100) [$\text{M} + \text{H}$]⁺. HRMS (ESI): Found 192.0619. Calc. for $\text{C}_{11}\text{H}_8\text{F}_2\text{N}$: 192.0619. The NMR data are in accordance with those reported in the literature.¹¹⁴

6-Methoxy-2-(4-methylphenyl)pyridine (1m). The general procedure was followed. 2-Bromo-6-methoxypyridine (0.803 g, 4.27 mmol, 1.0 equiv.), 4-methylphenylboronic acid (0.638 g, 4.69 mmol, 1.1 equiv.), K_3PO_4 (2.12 g, 10.0 mmol, 2.3 equiv.), $\text{Pd}(\text{OAc})_2$ (0.0187 g, 0.0883 mmol, 2.0 mol%) and PPh_3 (0.0664 g, 0.253 mmol, 5.9 mol%) were used. The obtained crude product was purified by flash column chromatography (98% hexanes/2% EtOAc), furnishing **1m** as a colourless oil (0.697 g, 3.50 mmol, 82%). ^1H NMR (400 MHz, CDCl_3): δ 7.95 (d, $^3J_{\text{H,H}} = 8.2$ Hz, 2H, H²), 7.59–7.63 (m, 1H, H⁴), 7.32 (d, $^3J_{\text{H,H}} = 7.4$ Hz, 1H, H³), 7.25–7.28 (m (partially overlap with the resonance corresponding to residual CHCl_3), 2H, H^{3'}), 6.66 (d, $^3J_{\text{H,H}} = 8.2$ Hz, 1H, H⁵), 4.04 (s, 3H, OCH_3), 2.41 ppm (s, 3H, Ar- CH_3). ^{13}C NMR (101 MHz, CDCl_3): δ 163.6 (C⁶), 154.7 (C²), 139.07 (C⁴), 138.77 (C⁴), 136.3 (C¹), 129.3 (C³), 126.6 (C²), 112.4 (C³), 108.8 (C⁵), 53.1 (OCH_3), 21.2 ppm (Ar- CH_3). ^{15}N (^1H) NMR (600 MHz, CD_2Cl_2): δ -121.5 ppm (N¹). MS (ESI): *m/z* (rel. %): 222.089 (100) [$\text{M} + \text{Na}$]⁺. HRMS (ESI): Found 222.0887. Calc. for $\text{C}_{13}\text{H}_{13}\text{NNaO}$: 222.0889. The NMR data are in accordance with those reported in the literature.¹¹⁵



2-(2,2'',3,3'',4,4'',5,5'',6,6''-Decafluoro-[1,1':3',1''-terphenyl]-5'-yl)pyridine (**1r**). The synthesis of **1r** was adapted from a literature procedure.⁷¹ K₂CO₃ (1.11 g, 8.00 mmol, 4.0 equiv.), Pd(OAc)₂ (0.0225 g, 0.100 mmol, 5.0 mol%) and S-Phos (0.0821 g, 0.200 mmol, 10 mol%) were added to a Schlenk flask, and flushed with Ar for 15 min. Pentafluorobenzene (1.00 g, 6.00 mmol, 3.0 equiv.) and *i*-PrOAc (4 mL) were added, followed by a solution of **16** (0.448 g, 2.00 mmol, 1.0 equiv.) in *i*-PrOAc (2 mL). After flushing with Ar, the reaction flask was sealed and stirred at 80 °C for 14 h. After cooling to rt, insoluble material was separated from the reaction mixture by filtration. The solids were washed with additional *i*-PrOAc, and the washings were combined with the original filtrate. The solvent was removed under reduced pressure, and the obtained solid was purified by flash column chromatography (85% hexanes/15% EtOAc), followed by two-fold recrystallization from EtOH. **1r** was obtained as a colourless fluffy solid (0.356 g, 0.730 mmol, 37%). ¹H NMR (600 MHz, CDCl₃): δ 8.72–8.73 (m, 1H, H⁶), 8.18 (s, 2H, H²), 7.78–7.83 (m, 2H, H⁴ + H³), 7.54 (s, 1H, H⁴), 7.31 ppm (ddd, ³J_{H,H} = 7.8 Hz, ³J_{H,H} = 4.8 Hz, ⁴J_{H,H} = 1.4 Hz, 1H, H⁵). ¹³C NMR (201 MHz, CDCl₃): δ 155.5 (C²), 150.0 (C⁶), 144.3 (d, ¹J_{C,F} = 248.5 Hz, Ar–C–F), 140.80 (d, ¹J_{C,F} = 254.8 Hz, Ar–C–F), 140.74 (C¹), 137.9 (d, ¹J_{C,F} = 251.3 Hz, Ar–C–F), 137.1 (C⁴), 132.0 (C⁴), 129.5 (C²), 127.6 (C³), 123.0 (C⁵), 120.6 (C³), 114.87–115.05 ppm (m, Ar–C–C–F). ¹⁹F NMR (376 MHz, CDCl₃): δ –145.6 (dd, ³J_{F,F} = 22.6 Hz, ⁴J_{F,F} = 6.0 Hz, 4F, C₆F₅), –157.4 (dd, ³J_{F,F} = 21.0 Hz, ³J_{F,F} = 20.9 Hz, 2F, C₆F₅), –164.6 ppm (m, 4F, C₆F₅). ¹⁵N{¹H} NMR (600 MHz, CD₂Cl₂): δ –73.0 ppm (N¹). MS (ESI): *m/z* (rel. %): 510.031 (100) [M + Na]⁺. HRMS (ESI): Found 510.0311. Calc. for C₂₃H₇F₁₀NNa: 510.0311. Elemental analysis: Found: C, 56.7; H, 1.4; N, 2.9. Calc. for C₂₃H₇F₁₀N: C, 56.7; H, 1.45; N, 2.9%.

2-(3,5-Diisopropylphenyl)pyridine (**1s**). The general procedure was followed. 2-Bromopyridine (0.924 g, 5.85 mmol, 1.0 equiv.), 3,5-diisopropylphenylboronic acid (1.20 g, 5.84 mmol, 1.0 equiv.), K₃PO₄ (2.48 g, 11.7 mmol, 2.0 equiv.), Pd(OAc)₂ (0.0273 g, 0.122 mmol, 2.1 mol%) and PPh₃ (0.0921 g, 0.352 mmol, 6.0 mol%) were used. The crude product was purified twice by flash column chromatography (first 85% hexanes/15% EtOAc, then CH₂Cl₂), yielding **1s** as a colourless oil (0.847 g, 3.54 mmol, 61%). ¹H NMR (600 MHz, CDCl₃): δ 8.70 (d, ³J_{H,H} = 4.8 Hz, 1H, H⁶), 7.72–7.75 (m, 2H, H³ + H⁴), 7.66 (s, 2H, H²), 7.21 (ddd, ³J_{H,H} = 6.1 Hz, ³J_{H,H} = 4.8 Hz, ⁴J_{H,H} = 2.5 Hz, 1H, H⁵), 7.15 (s, 1H, H⁴), 2.99 (sp, ³J_{H,H} = 7.0 Hz, 2H, CH(CH₃)₂), 1.31 ppm (d, ³J_{H,H} = 7.0 Hz, 12H, CH(CH₃)₂). ¹³C NMR (151 MHz, CDCl₃): δ 158.2 (C²), 149.6 (C⁶), 149.3 (C³), 139.4 (C¹), 136.6 (C⁴), 125.4 (C⁴), 122.7 (C²), 121.8 (C⁵), 120.8 (C³), 34.3 (CH(CH₃)₂), 24.1 ppm (CH(CH₃)₂). ¹⁵N{¹H} NMR (600 MHz, CD₂Cl₂): δ –73.6 ppm (N¹). MS (ESI): *m/z* (rel. %): 240.175 (100) [M + H]⁺. HRMS (ESI): Found 240.1746. Calc. for C₁₇H₂₂N [M + H]⁺: 240.1747. Elemental analysis: Found: C, 85.3; H, 8.9; N, 5.85. Calc. for C₁₇H₂₁N: C, 85.3; H, 8.8; N, 5.85%.

2-(3,5-Diethylphenyl)pyridine (**1t**). The general procedure was followed. 2-Bromopyridine (0.670 g, 4.26 mmol, 1.0 equiv.), 3,5-diethylphenylboronic acid (0.670 g, 4.26 mmol, 1.0

equiv.), K₃PO₄ (1.81 g, 8.52 mmol, 2.0 equiv.), Pd(OAc)₂ (0.0187 g, 0.0831 mmol, 2.0 mol%) and PPh₃ (0.0664 g, 0.253 mmol, 6.0 mol%) were used. The crude product was purified by flash column chromatography (85% hexanes/15% EtOAc), yielding **1t** as a colourless oil (0.438 g, 2.070 mmol, 49%). ¹H NMR (600 MHz, CDCl₃): δ 8.69–8.71 (ddd, ³J_{H,H} = 4.7 Hz, ⁴J_{H,H} = 1.3 Hz, ⁵J_{H,H} = 1.3 Hz, 1H, H⁶), 7.72–7.74 (m, 2H, H³ + H⁴), 7.65 (d, ⁴J_{H,H} = 1.2 Hz, 2H, H²), 7.20–7.22 (m, 1H, H⁵), 7.11 (s, 1H, H⁴), 2.73 (q, ³J_{H,H} = 7.6 Hz, 4H, CH₂CH₃), 1.30 ppm (t, ³J_{H,H} = 7.6 Hz, 6H, CH₂CH₃). ¹³C NMR (151 MHz, CDCl₃): δ 158.0 (C²), 149.5 (C⁶), 144.7 (C³), 139.4 (C¹), 136.6 (C⁴), 128.3 (C⁴), 123.9 (C²), 121.8 (C⁵), 120.7 (C³), 28.6 (CH₂CH₃), 15.6 ppm (CH₂CH₃). ¹⁵N{¹H} NMR (600 MHz, CD₂Cl₂): δ –73.6 ppm (N¹). MS (ESI): *m/z* (rel. %): 212.143 (100) [M + H]⁺. HRMS (ESI): Found 212.1433. Calc. for C₁₅H₁₈N [M + H]⁺: 212.1434. Elemental analysis: Found: C, 85.2; H, 8.1; N, 6.6. Calc. for C₁₅H₁₇N: C, 85.3; H, 8.1; N, 6.6%.

2c-Au(OAc)^F₂. A microwave vessel was charged with Au(OAc)₃ (0.373 g, 0.999 mmol, 1.00 equiv.) and **1c** (0.193 g, 1.01 mmol, 1.01 equiv.) in a 1 : 1 mixture of HOAc^F and water (30 mL). The reaction mixture was heated at 120 °C for 30 min in a microwave. After cooling to room temperature, HOAc^F (30 mL) was added to dissolve partially precipitated product, and the resulting solution was filtered. Water (50 mL) was added to the filtrate, resulting in the precipitation of a solid. After cooling on an ice-water bath for 15 min, the precipitate was filtered off, washed with water (3 × 5 mL) and Et₂O (5 mL), and dried under a stream of air for ca. 3 h, furnishing 2c-Au(OAc)^F₂ as a colourless solid (0.527 g, 0.860 mmol, 85%). ¹H NMR (800 MHz, CD₂Cl₂): δ 8.62 (dd, ³J_{H,H} = 6.0 Hz, ⁴J_{H,H} = 1.0 Hz, H⁶), 8.33 (ddd, ³J_{H,H} = 7.9 Hz, ³J_{H,H} = 7.8 Hz, ⁴J_{H,H} = 1.4 Hz, H⁴), 7.95 (d, ³J_{H,H} = 8.1 Hz, H³), 7.66 (ddd, ³J_{H,H} = 7.6 Hz, ³J_{H,H} = 6.0 Hz, ⁴J_{H,H} = 1.3 Hz, H⁵), 7.28 (dd, ³J_{H,F} = 7.8 Hz, ⁴J_{H,H} = 2.4 Hz, H⁶), 6.87 ppm (ddd, ³J_{H,F} = 9.1 Hz, ³J_{H,F} = 9.1 Hz, ⁴J_{H,H} = 2.5 Hz, H⁴). ¹³C NMR (201 MHz, CD₂Cl₂): δ 164.1 (C²), 164.0 (dd, ¹J_{C,F} = 251.5 Hz, ³J_{C,F} = 12.0 Hz, C³ or C⁵), 162.8 (dd, ¹J_{C,F} = 253.4 Hz, ³J_{C,F} = 12.1 Hz, C³ or C⁵), 161.1 (q, ²J_{C,F} = 38.2 Hz, OCOCF₃), 161.0 (q, ²J_{C,F} = 39.5 Hz, OCOCF₃), 148.2 (C⁶), 145.4 (C⁴), 145.1 (dd, ³J_{C,F} = 9.9 Hz, ³J_{C,F} = 9.6 Hz, C¹), 126.4 (C⁵), 122.9 (C³), 118.8 (dd, ²J_{C,F} = 25.3 Hz, ⁴J_{C,F} = 3.6 Hz, C²), 118.0 (q, ¹J_{C,F} = 288.6 Hz, OCOCF₃), 115.5 (q, ¹J_{C,F} = 287.8 Hz, OCOCF₃), 109.2–109.5 ppm (m, C⁴ + C⁶). ¹⁹F NMR (188 MHz, CD₂Cl₂): δ –75.5 (d, ⁷J_{F,F} = 4.0 Hz, 3F, OCOCF₃ (*trans*-N)), –77.0 (s, 3F, OCOCF₃ (*cis*-N)), –102.46 to –102.53 (m, 1F, ArF¹), –109.2 ppm (d, ⁴J_{F,F} = 10.0 Hz, 1F, ArF²). MS (ESI): *m/z* (rel. %): 499.998 (90) [M – OCOCF₃]⁺. HRMS (ESI): Found: 499.9979. Calc. for C₁₃H₆AuF₅NO₂ [M – OCOCF₃]⁺: 499.9979. Elemental analysis: Found: C, 29.4; H, 1.0; N, 2.3. Calc. for C₁₅H₆AuF₈NO₄: C, 29.4; H, 1.0; N, 2.3%.

2m-Au(OAc)^F₂. A microwave vessel was charged with Au(OAc)₃ (0.0750 g, 0.200 mmol, 1.00 equiv.) and **1m** (0.0400 g, 0.202 mmol, 1.01 equiv.) in a 1 : 1 mixture of HOAc^F and water (6 mL). The reaction mixture was heated at 100 °C for 60 min in a microwave. After cooling to room temperature, HOAc^F (1 mL) was added, and the resulting solution was filtered. Water (8 mL) was added to the filtrate, resulting in the precipi-



tation of a solid. After cooling on an ice-water bath for 10 min, the precipitate was filtered off, washed with water (3 × 5 mL) and Et₂O (5 mL), and dried under a stream of air for ca. 3 h, furnishing **2m-Au(OAc^F)₂** as a pale yellow solid (0.0660 g, 0.106 mmol, 53%). ¹H NMR (800 MHz, CD₂Cl₂): δ 8.09 (dd, ³J_{H,H} = 8.2 Hz, ³J_{H,H} = 7.8 Hz, 1H, H⁴), 7.42 (dd, ³J_{H,H} = 7.7 Hz, ⁴J_{H,H} = 0.9 Hz, 1H, H³), 7.35 (d, ³J_{H,H} = 7.9 Hz, 1H, H⁶), 7.23–7.25 (m, 1H, H⁵), 6.82 (d, ⁴J_{H,H} = 0.6 Hz, 1H, H³), 6.81 (dd, ³J_{H,H} = 8.5 Hz, ⁴J_{H,H} = 0.8 Hz, 1H, H⁵), 4.01 (s, 3H, OCH₃), 2.40 ppm (s, 3H, Ar-CH₃). ¹³C NMR (201 MHz, CD₂Cl₂): δ 165.1 (C⁶), 163.7 (C²), 161.1 (q, ²J_{C,F} = 36.6 Hz, OCOCF₃), 160.4 (q, ²J_{C,F} = 39.6 Hz, OCOCF₃), 146.6 (C⁴), 143.8 (C⁴), 141.7 (C²), 139.5 (C¹), 131.3 (C⁵), 128.7 (C³), 125.9 (C⁶), 118.7 (q, ¹J_{C,F} = 289.6 Hz, OCOCF₃), 116.1 (q, ¹J_{C,F} = 288.1 Hz, OCOCF₃), 113.3 (C³), 106.9 (C⁵), 58.0 (OCH₃), 22.3 ppm (Ar-CH₃). ¹⁹F NMR (376 MHz, CD₂Cl₂): δ -76.2 (s, 3F, OCOCF₃), -76.5 ppm (broadened s, 3F, OCOCF₃). ¹⁵N{¹H} NMR (600 MHz, CD₂Cl₂): δ -187.7 ppm (N¹). MS (ESI): *m/z* (rel. %): 426.076 (100) [M - 2OCOCF₃ + OMe]⁺. HRMS (ESI): Found: 426.0761. Calc. for C₁₄H₁₅AuNO₂ [M - 2OCOCF₃ + OMe]⁺: 426.0763. Elemental analysis: Found: C, 32.8; H, 2.0; N, 2.3. Calc. for C₁₇H₁₂AuF₆NO₅: C, 32.9; H, 1.95; N, 2.25%.

2r-Au(OAc^F)₂. A microwave vessel was charged with Au(OAc)₃ (0.0750 g, 0.200 mmol, 1.00 equiv.) and **1r** (0.0980 g, 0.202 mmol, 1.01 equiv.) in a 1 : 1 mixture of HOAc^F and water (6 mL). The reaction mixture was heated at 120 °C for 60 min in a microwave. After cooling to room temperature, HOAc^F (2 mL) was added, and the resulting solution was filtered. Water (20 mL) was added to the filtrate, resulting in the precipitation of a solid. After cooling on an ice-water bath for 15 min, and then overnight at 4–8 °C, the precipitate was filtered off, washed with water (3 × 5 mL) and Et₂O (5 mL), and dried under a stream of air for ca. 3 h, furnishing **2r-Au(OAc^F)₂** as a colourless solid (0.0740 g, 0.0810 mmol, 41%). ¹H NMR (800 MHz, CD₂Cl₂): δ 8.61 (dd, ³J_{H,H} = 6.1 Hz, ⁴J_{H,H} = 1.0 Hz, 1H, H⁶), 8.34 (ddd, ³J_{H,H} = 8.1 Hz, ³J_{H,H} = 7.7 Hz, ⁴J_{H,H} = 1.4 Hz, 1H, H⁴), 8.05 (dd, ³J_{H,H} = 8.2 Hz, ⁴J_{H,H} = 1.2 Hz, 1H, H³), 7.80 (s, 1H, H⁶), 7.65 (ddd, ³J_{H,H} = 7.6 Hz, ³J_{H,H} = 6.1 Hz, ⁴J_{H,H} = 1.4 Hz, 1H, H⁵), 7.27 ppm (s, 1H, H⁴). ¹³C NMR (201 MHz, CD₂Cl₂): δ 164.3 (C²), 161.2 (q, ²J_{C,F} = 38.4 Hz, OCOCF₃), 160.4 (q, ²J_{C,F} = 39.6 Hz, OCOCF₃), 148.4 (C⁶), 145.4 (C⁴), 144.6 (m, 2 × Ar-C-F), 142.2 (d, ¹J_{C,F} = 254.8 Hz, Ar-C-F), 141.9 (d, ¹J_{C,F} = 256.2 Hz, Ar-C-F), 141.1 (C³ + C⁵), 138.6 (d, ¹J_{C,F} = 252.5 Hz, Ar-C-F), 137.6 (C⁴), 137.2 (d, ¹J_{C,F} = 251.1 Hz, Ar-C-F), 130.8 (C¹ or C²), 129.4 (C¹ or C²), 128.3 (C⁶), 126.3 (C⁵), 122.8 (C³), 117.9 (q, ¹J_{C,F} = 288.4 Hz, OCOCF₃), 115.1 (q, ¹J_{C,F} = 288.1 Hz, OCOCF₃), 113.0 ppm (m, 2 × Ar-C-C-F). ¹⁹F NMR (470 MHz, CD₂Cl₂): δ -75.6 (s, 3F, OCOCF₃), -77.1 (s, 3F, OCOCF₃), -143.8 (dd, ³J_{F,F} = 22.1 Hz, ⁴J_{F,F} = 7.5 Hz, 2F, C₆F₅), -145.2 (dd, ³J_{F,F} = 23.1 Hz, ⁴J_{F,F} = 7.5 Hz, 2F, C₆F₅), -155.0 (dd, ³J_{F,F} = 21.3 Hz, ³J_{F,F} = 20.3 Hz, 1F, C₆F₅), -156.3 (dd, ³J_{F,F} = 20.8 Hz, ³J_{F,F} = 20.2 Hz, 1F, C₆F₅), -163.6 (m, 2F, C₆F₅), -164.7 ppm (m, 2F, C₆F₅). ¹⁵N{¹H} NMR (800 MHz, CD₂Cl₂): δ -167.7 ppm (N¹). MS (ESI): *m/z* (rel. %): 714.019 (15) [M - 2OCOCF₃ + OMe]⁺, 768.027 (70) [M - 2OCOCF₃ + 2OMe + Na]⁺, 800.053 (100) [M - 2OCOCF₃ + 2OMe + MeOH + Na]⁺. HRMS (ESI): Found:

768.0269. Calc. for C₂₅H₁₂AuF₁₀NNaO₂ [M - 2OCOCF₃ + 2OMe + Na]⁺: 768.0266. Elemental analysis: Found: C, 35.6; H, 0.7; N, 1.5. Calc. for C₂₆H₆AuF₁₆NO₄: C, 35.7; H, 0.7; N, 1.5%.

3s-AuOAc^F. A microwave vessel was charged with Au(OAc)₃ (0.0750 g, 0.200 mmol, 1.00 equiv.) and **1s** (0.0480 g, 0.202 mmol, 1.01 equiv.) in a 1 : 1 mixture of HOAc^F and water (6 mL). The reaction mixture was heated at 80 °C for 3.5 h in a microwave. The reaction mixture was kept at 4–8 °C overnight. Addition of HOAc^F (4 mL), followed by water (6 mL), furnishing a white precipitate. The precipitate was collected by filtration, and washed with water (3 × 3 mL). The title compound was obtained as a colourless solid (0.0410 g, 0.0750 mmol, 38%). ¹H NMR (600 MHz, CD₂Cl₂): δ 8.48 (ddd, ³J_{H,H} = 5.4 Hz, ⁴J_{H,H} = 1.4 Hz, ⁵J_{H,H} = 0.8 Hz, 1H, H⁶), 7.99–8.03 (m, 1H, H⁴), 7.94 (d, ³J_{H,H} = 8.1 Hz, 1H, H³), 7.50 (ddd, ³J_{H,H} = 7.5 Hz, ³J_{H,H} = 5.4 Hz, ⁴J_{H,H} = 1.3 Hz, 1H, H⁵), 7.35 (s, 1H, H⁶), 6.92 (s, 1H, H⁴), 3.50 (sx, ³J_{H,H} = 6.8 Hz, 1H, Ar-CH(CH₂Au)CH₃), 3.38 (dd, ²J_{H,H} = 10.4 Hz, ³J_{H,H} = 7.3 Hz, 1H, Ar-CH(CH₂Au)CH₃), 2.97 (dd, ²J_{H,H} = 10.4 Hz, ³J_{H,H} = 5.6 Hz, 1H, Ar-CH(CH₂Au)CH₃), 2.94 (sp, ³J_{H,H} = 6.9 Hz, 1H, Ar-CH(CH₃)₂), 1.35 (d, ³J_{H,H} = 7.0 Hz, 3H, Ar-CH(CH₂Au)CH₃), 1.290 (d, ³J_{H,H} = 6.9 Hz, 3H, Ar-CH(CH₃)₂), 1.287 ppm (d, ³J_{H,H} = 6.9 Hz, 3H, Ar-CH(CH₃)₂). ¹³C NMR (151 MHz, CD₂Cl₂): δ 162.2 (C²), 161.5 (q, ²J_{C,F} = 36.5 Hz, OCOCF₃), 161.2 (C³), 149.8 (C⁵), 148.1 (C⁶), 146.3 (C¹), 141.5 (C⁴), 139.7 (C²), 126.1 (C⁴), 125.1 (C⁵), 121.1 (C⁶), 120.8 (C³), 118.5 (q, ¹J_{C,F} = 290.3 Hz, OCOCF₃), 46.7 (Ar-CH(CH₂Au)CH₃), 41.4 (Ar-CH(CH₂Au)CH₃), 35.2 (Ar-CH(CH₃)₂), 24.5 (Ar-CH(CH₃)₂), 24.3 (Ar-CH(CH₃)₂), 23.0 ppm (Ar-CH(CH₂Au)CH₃). ¹⁹F NMR (376 MHz, CD₂Cl₂): δ -76.9 ppm (s, 3F, OCOCF₃). ¹⁵N{¹H} NMR (600 MHz, CD₂Cl₂): δ -111.7 ppm (N¹). MS (ESI): *m/z* (rel. %): 434.118 (49) [M - OCOCF₃]⁺, 452.128 (100) [M - OCOCF₃ + H₂O]⁺, 475.144 (82) [M - OCOCF₃ + MeCN]⁺. HRMS (ESI): Found: 434.1176. Calc. for C₁₇H₁₉AuN [M - OCOCF₃]⁺: 434.1178. Elemental analysis: Found: C, 41.6; H, 3.5; N, 2.5. Calc. for C₁₇H₁₅AuF₃NO₂: C, 41.7; H, 3.5; N, 2.6%.

3t-AuOAc^F. A microwave vessel was charged with Au(OAc)₃ (0.0750 g, 0.200 mmol, 1.00 equiv.) and **1t** (0.0440 g, 0.202 mmol, 1.04 equiv.) in a 1 : 1 mixture of HOAc^F and water (6 mL). The reaction mixture was heated at 80 °C for 3 h in a microwave. The reaction mixture was kept at 4–8 °C overnight. Addition of HOAc^F (2 mL), followed by water (3 mL), furnishing a white precipitate. The precipitate was collected by filtration, and washed with water (3 × 3 mL) and pentane (5 mL). The title compound was obtained as a colourless solid (0.0680 g, 0.131 mmol, 65%). ¹H NMR (600 MHz, CD₂Cl₂): δ 8.45–8.47 (m, 1H, H⁶), 7.99–8.01 (m, 1H, H⁴), 7.90 (d, ³J_{H,H} = 8.0 Hz, 1H, H³), 7.47 (ddd, ³J_{H,H} = 7.6, ³J_{H,H} = 5.4, ⁴J_{H,H} = 1.3 Hz, 1H, H⁵), 7.31 (s, 1H, H⁶), 6.99 (s, 1H, H⁴), 3.27–3.30 (m, 2H, Ar-CH₂CH₂Au), 3.18–3.20 (m, 2H, Ar-CH₂CH₂Au), 2.65 (q, ³J_{H,H} = 7.6 Hz, 2H, Ar-CH₂CH₃), 1.25 ppm (t, ³J_{H,H} = 7.6 Hz, 3H, Ar-CH₂CH₃). ¹³C NMR (151 MHz, CD₂Cl₂): δ 162.5 (C²), 161.5 (d (q expected), ²J_{C,F} = 36.3 Hz, OCOCF₃), 158.3 (C³), 148.2 (C⁵), 148.0 (C⁶), 145.0 (C¹), 141.5 (C⁴), 139.4 (C²), 126.9 (C⁴), 125.0 (C⁵), 122.0 (C⁶), 120.9 (C³), 118.5 (d (q expected), ¹J_{C,F} = 290.5 Hz, OCOCF₃), 40.3 (Ar-CH₂CH₂Au), 32.2 (Ar-



CH₂CH₂Au), 29.7 (Ar-CH₂CH₃), 16.2 ppm (Ar-CH₂CH₃). ¹⁹F NMR (376 MHz, CD₂Cl₂): δ -76.9 ppm (s, 3F, OCOCF₃). ¹⁵N {¹H} NMR (600 MHz, CD₂Cl₂): δ -110.7 ppm (N¹). MS (ESI): *m/z* (rel. %): 406.087 (63) [M - OCOCF₃]⁺, 424.097 (100) [M - OCOCF₃ + H₂O]⁺, 447.113 (54) [M - OCOCF₃ + MeCN]⁺. HRMS (ESI): Found: 406.0867. Calc. for C₁₅H₁₅AuN [M - OCOCF₃]⁺: 406.0865. Elemental analysis: Found: C, 39.3; H, 2.9; N, 2.7. Calc. for C₁₇H₁₅AuF₃NO₂: C, 39.3; H, 2.9; N, 2.7%.

Computational details

Calculations were carried out at the DFT level as implemented in the Gaussian16 software package.¹¹⁶ The hybrid PBE0+GD3 functional^{117,118} including Grimme's model for dispersion forces was used to optimize all geometries. This methodology was selected based on previous studies which have proven its solid performance in the modelling of Au(III) complexes.^{7,24,25,119,120} C, H, F, N and O were described with the all-electron triple- ζ 6-311+G** basis set,^{121,122} whereas Au was described with the Stuttgart-Köln basis set including a small-core quasi-relativistic pseudopotential.^{123,124} NBO7 calculations were performed in order to analyse the natural charges.¹²⁵ Geometries were fully optimized without any constraint. Vibrational frequencies were computed at the same level of theory to classify all stationary points as either saddle points (transition states, with a single imaginary frequency) or energy minima (reactants, intermediates and products, with only real frequencies). The Gibbs free energy used in the discussion includes both the thermochemistry and the refined energy. All optimizations were carried out in solvent (CH₂Cl₂ or HOAc^F) using the SMD solvation model.¹²⁶ HOAc^F was defined as eps = 8.55, epsinf = 2.26 and rsolv = 13.7. In the bimolecular steps, the energies were corrected for the 1 M standard state.

Author contributions

KTH: investigation, supervision, writing – original draft, writing – review and editing; ILS: investigation, writing – review and editing; DSW: investigation, supervision; AN: supervision, writing – review and editing; MT: supervision, writing – review and editing.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the Research Council of Norway through the Norwegian NMR Package in 1994, through the Norwegian NMR Platform, NNP (226244/F50) and through the Hylleraas Centre for Quantum Molecular Sciences (project number 262695). This work was also supported by the ERASMUS program of the European Union (exchange visit to

Oslo for ILS), and the Norwegian Metacenter for Computational Science (NOTUR, nn4654k). Dr Marte S. M. Holmsen, Dr Sigurd Øien-Ødegaard, Lorena P. Escrivá, Michael Philipp and Sahra A. Ahmed are acknowledged for experimental help. We thank Osamu Sekiguchi, Lina Aarsbog and Sverre Løyland for performing the MS experiments, and Dr Richard H. Heyn (SINTEF Industry) for assistance with the elemental analyses. We thank Prof. Odile Eisenstein for helpful discussions. We acknowledge use of the Norwegian National Centre for X-ray Diffraction and Scattering (RECX).

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